

## Inhibitive effect by *Psidium guajava* leaf extract on the corrosion of double thermally-aged Al-Si-Mg (SSM-HPDC) alloy in simulated seawater environment

L. M. Sani,<sup>a,\*</sup> M. Abdulwahab,<sup>a</sup> S. A. Yaro<sup>a,b</sup> and F. O. Kolawole<sup>c</sup>

<sup>a</sup>Department of Metallurgical and Materials Engineering, Ahmadu Bello University, Zaria, Nigeria.

<sup>b</sup>Shell chair professorial occupant, Dept of Mechanical Engineering, Ahmadu Bello Univ., Zaria, Nigeria.

<sup>c</sup>Department of Materials and Metallurgical Engineering, Federal University Oye-Ekiti, Nigeria.

Received 15 April 2015; accepted 04 March 2016

---

### Abstract

The assessment of *Psidium guajava* leaf extract as corrosion inhibitor for double thermally-aged Al-Si-Mg (SSM-HPDC) alloy in 3.5%wt NaCl solution using the gravimetric and potentiodynamic polarization techniques was investigated. The gravimetric test was carried out at different inhibitor concentration, time and temperature ranges of 0.1-0.5%v/v, 1-5 hrs and 30-70°C. The results revealed that *Psidium guajava* leaf extract in 3.5%wt NaCl solution-aluminium environment decreased the corrosion rate at various concentrations considered. Inhibition efficiency (IE) of 72.1% at 0.5% v/v *Psidium guajava* leaf extract addition using the gravimetric method was demonstrated in 3.5%wt NaCl solution. The IE from the potentiodynamic polarization method was significantly enhanced as high as 98.89%/0.5%v/v. The additions of *Psidium guajava* leaf extract as corrosion inhibitor in the solution indicate higher potential value, IE and polarization resistance with decrease in current density. The methods for assessment of the alloy were in agreement and mixed-type corrosion exist which obeyed the Langmuir adsorption isotherms. The optical microscopy (OPM) revealed that the microstructure of the double thermally-aged sample has finer grains and enhanced grain boundaries compared to the untreated sample. The scanning electron microscope (SEM) surface morphology of as-corroded uninhibited condition showed severe damage and pit formation than as-corroded inhibited.

**Keywords:** Aluminium, *Psidium guajava*, Corrosion, Thermal ageing, SEM, OPM, Thin film.

---

### Introduction

Aluminum is the second widely used metal due to its desirable chemical, physical and mechanical properties. It is alloyed with elements like Si, Mg, Cu,

---

\* Corresponding author. E-mail address: sanilawal66@yahoo.com

Mn, Fe etc. Addition of silicon to aluminum can improve its fluidity as well as castability and mechanical properties [2]. Addition of Magnesium to Al-Si alloy improves its strength to weight ratio and yield stress by combining with silicon to form the age hardening phase ( $Mg_2Si$ ) which precipitates from a super saturated solid solution during heat treatment [4].

Aluminium and its alloys are widely used for different applications in industries and marine environment because of their excellent properties [12]. However, the alloys still exhibit insufficient and poor performance in service condition that requires high hardness, wear and aggressive corrosion environment [1]. Due to different industrial applications and economic importance of aluminium and its alloys, their protection against corrosion by inhibiting the working environment is worth studying [3, 10]. The safety and environmental issues of corrosion inhibitors arising in industries has always been a global concern [13]. The use of inhibitors for corrosion control of materials, which are in contact with aggressive environment, is an accepted practice [11]. The corrosion resistance of aluminium and its alloys using inhibitors have been investigated in many environments [5, 7, 8, 10, 14, 15]. Although synthetic inhibitors when used for corrosion control, have been reported to indicate an excellent performance. But majority of these inhibitors are not eco-friendly and are expensive [12]. Therefore effort towards identifying any potential eco-friendly and less expensive corrosion inhibitors remain relevant and important. The growing interest among researchers for green inhibitors remained a top research focus. In this direction, plant extracts and oil have gained acceptance as corrosion inhibitors that are considered safe, eco - friendly, available and less expensive for most metals and alloys [12]. Accordingly, the present study involved the use of guava (*Psidium guajava*) leaf extracts as corrosion inhibitors for double thermally-aged Al-Si-Mg (SSM-HPDC) alloy.

## Experimental Procedure

### *Materials and sample preparation*

Aluminium alloy specimen with chemical composition as shown in Table 1 was sectioned into coupons of dimension 12×12×2 mm for the corrosion study in *Psidium guajava*- 3.5%wt NaCl solution.

**Table 1:** Chemical composition of aluminium alloy used

Element	Al	Si	Mg	Fe	Cu	Mn	Zn	Ti	Sr
%	92.28	7.14	0.36	0.10	0.01	0.01	0.01	0.07	0.02

Initially, the coupons were mechanically polished with emery papers down to 800 grit size. The coupons were degreased in ethanol, dried, weighed and stored in desiccators. The initial weight of each sample was taken and recorded.

### *Heat treatment*

The coupons were solution treated at a temperature of 540°C for 1hr in a heat treatment electric resistance furnace and then rapidly quenched in warm water

(65°C). The quenched samples were pre-aged at 105°C for 2hrs, and then finally aged at 180°C for 2 hrs before cooling in air. **Gravimetric measurement**

The corrosion rate by gravimetric analysis was carried out on the previous weighed coupons with and without inhibitor at 303K, 323K and 343K. The volume of the solution was 150ml with and without the addition of *Psidium guajava* inhibitor. The *Psidium guajava* inhibitor concentration was varied from 0.1, 0.2, 0.3, 0.5%v/v in 150ml of 3.5%wt NaCl solution. For each sample, using gravimetric method, the samples were washed, dried and weighed at interval of 1, 2, 3, 4 and 5 hrs of exposure time. The corrosion rate and inhibition efficiency were determined for each inhibitor concentration at 30, 50 and 70°C.

#### **Potentiodynamic polarization technique**

The potentiodynamic polarization and linear polarization resistance was used to characterize the corrosion rate of double thermally-aged aluminium alloy in *Psidium guajava*-simulated seawater medium. In the electrochemical test, a glass corrosion cell kit with a platinum counter electrode, a saturated Ag/Ag reference electrode and aluminium sample as working electrode were used. The working electrode samples were positioned at the glass corrosion cell kit, leaving 1.14cm<sup>2</sup> surfaces in contact with the solution. Polarization test was carried out in simulated seawater solution at room temperature using a potentiostat. The polarization curves were determined by stepping the potential at a scan rate of 0.003V/sec. The polarization curves were plotted using

Autolab data acquisition system and both the corrosion rate and potential were estimated by the Tafel extrapolation method.

#### **Microstructure and surface morphology**

The samples for the microscopic examination were ground with emery papers of progressively fine grade 80, 220, 600, 800 grits sizes using water as coolant. The ground samples were then polished using one-micro size alumina powder suspended in distilled water, followed by etching in Keller's reagent. The optical microscope (OPM) with an inbuilt camera was used for the microstructure observation of the as-received and age-hardened samples while Scanning electron microscope (SEM) was used to assess the surface morphology of the inhibited and uninhibited samples.

## **Results and Discussion**

#### **Hardness measurement and double thermal ageing treatment**

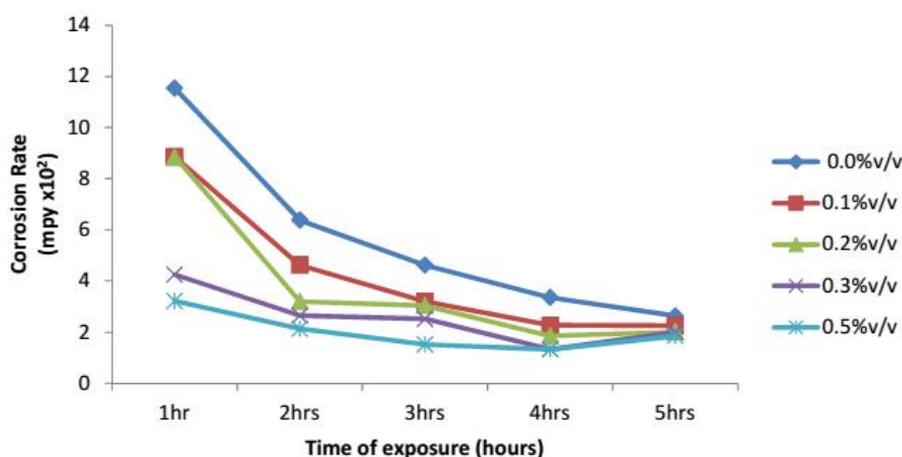
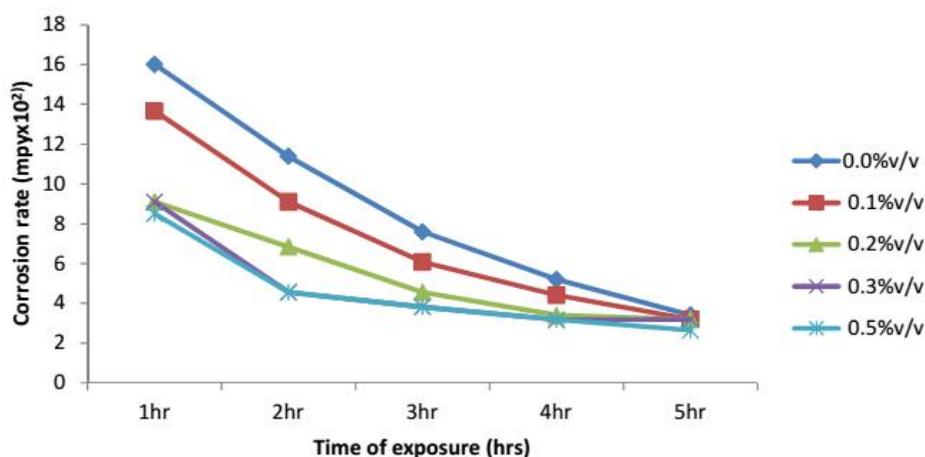
Table 2 presents the average hardness value for Al-Si-Mg (SSM-HPDC) alloy, the result of the age-hardened sample showed an increase in hardness value (82.8 HRF) compared to the as-received sample (59.1 HRF) indicating an increase of 28.62%. The age-hardening mechanisms responsible for strengthening are based on formation of intermetallic compounds during decomposition of metastable supersaturated solid solution obtained by solution treatment and quenching (Precipitation hardening)[9].

**Table 2:** Variation of hardness of Al-Si-Mg (SSM-HPDC) alloy in the as-received and age-hardened conditions.

Samples	1 <sup>st</sup> HRF	2 <sup>nd</sup> HRF	3 <sup>rd</sup> HRF	Mean HRF
As-received	61.0	58.5	57.8	59.1
Age-hardened	79.3	84.9	84.4	82.8

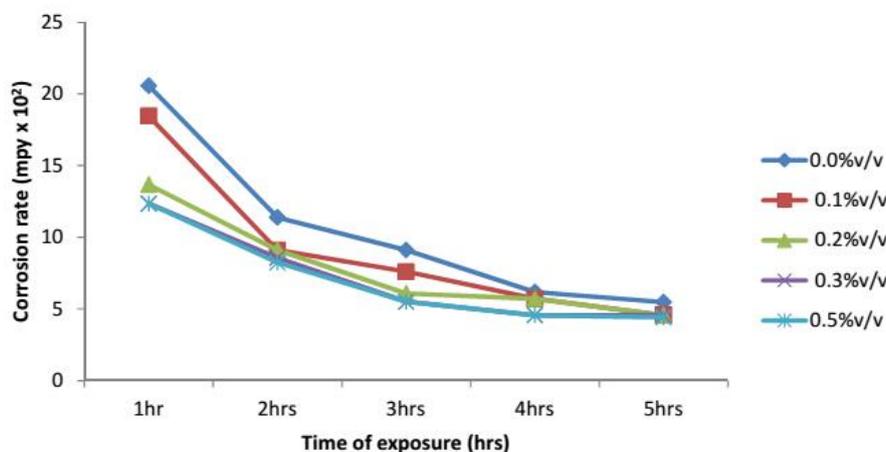
**Gravimetric-based mass loss measurement and corrosion rate**

From result, the corrosion rate (Figures 1-3) showed the variation of corrosion rate of the age-hardened alloy in simulated seawater environment decrease with addition of inhibitor. In Figures 1, the sample in NaCl/0.5%v/v *Psidium guajava* leaf extract exhibit an excellent corrosion resistance. Similar trend was observed in Figure 2 and 3. For example, corrosion rate decreased from  $11.54 \times 10^2$  mpy (control) to  $3.32 \times 10^2$  mpy (0.5%v/v) at 30°C for 1hr immersion time (Figure 1).

**Figure 1:** Variation of corrosion rate of double thermally aged Al-Si-Mg alloy with time of exposure in the presence of guava (*Psidium guajava*) leaf extract at 30°C.**Figure 2:** Variation of corrosion rate of double thermally aged Al-Si-Mg alloy with time of exposure in the presence of guava (*Psidium guajava*) leaf extract at 50°C.

While at 50°C/70°C (0.5%v/v) for 1hr immersion time, corrosion rate decreased from  $16.01 \times 10^2$  /  $20.58 \times 10^2$  mpy (control) to  $8.51 \times 10^2$  /  $12.35 \times 10^2$  mpy. The significant decrease in corrosion rate at 0.5%v/v inhibitor concentration can be attributed to the adsorption of inhibitor molecules on alloy surface which act as

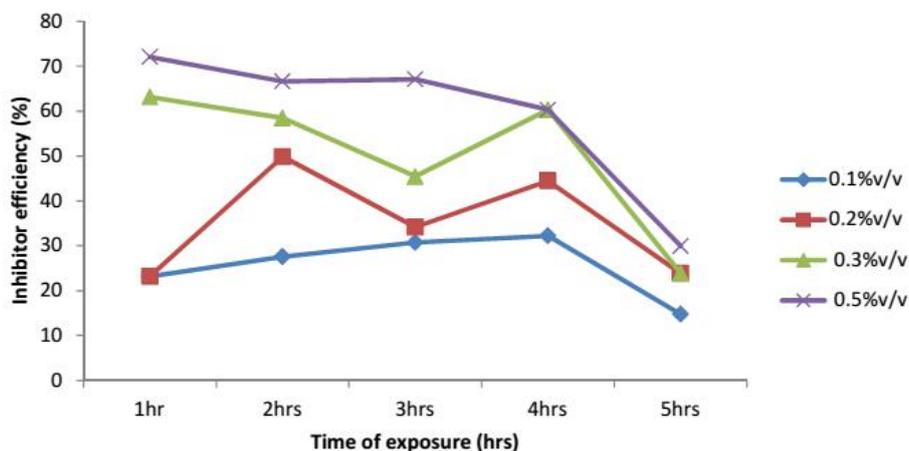
physical barrier to restrict the diffusion of ions to and from the alloy surface and prevent the alloy atoms (ions) from participating in further anodic or cathodic reactions, hence resulting in decrease in corrosion rate [3].



**Figure 3:** Variation of corrosion rate of double thermally aged Al-Si-Mg alloy with time of exposure in the presence of guava (*Psidium guajava*) leaf extract at 70°C.

#### ***Effect of inhibitor concentration and time on inhibitor efficiency.***

Figures 4-6 showed the variations of inhibitor efficiency with time for guava leaf extract with maximum inhibitor efficiencies of 72.10, 66.61, 67.10, 60.30, 29.92% for alloy with time of 1-5hrs respectively were obtained at concentration of 0.5%v/v at 30°C. This may be attributed to the fact that, as the extract addition increases, the surface area covered by the inhibitor increased hence higher IE was achieved [12].

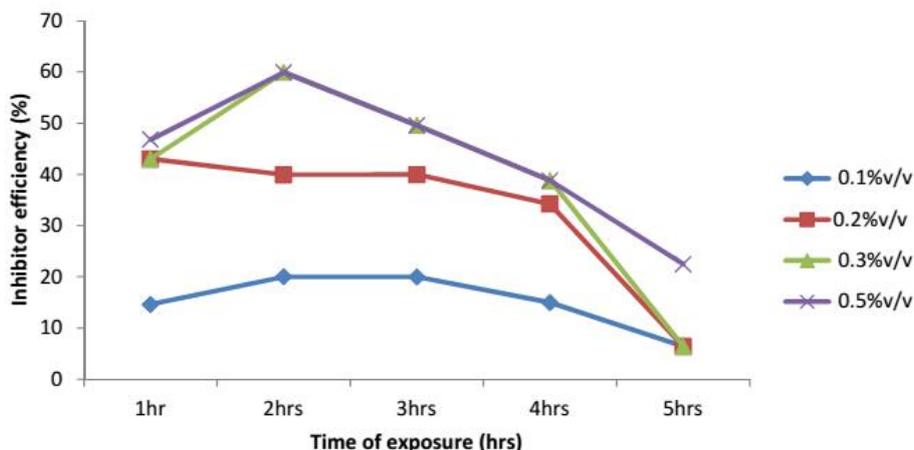


**Figure 4:** Variation of inhibitor efficiency of guava (*Psidium guajava*) leaf extract on double thermally aged Al-Si-Mg alloy with time of exposure at 30°C.

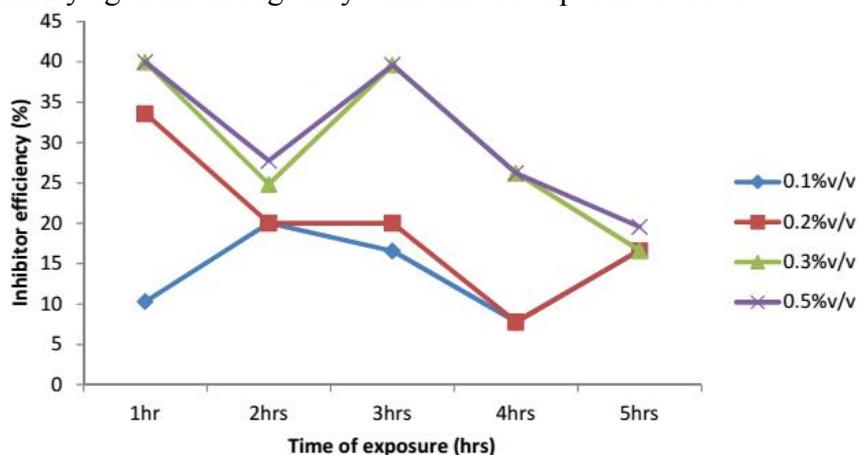
#### ***Effect of temperature on inhibitor efficiency***

It is evident from figure 7 that inhibition efficiency decreases with increasing temperature. The decrease in inhibitor efficiency with increase in temperature can be attributed to the fact that at a lower temperature the inhibitor molecules adsorbed onto the alloy surface, while at higher temperature increased rate of dissolution process takes place along with partial desorption of inhibitor

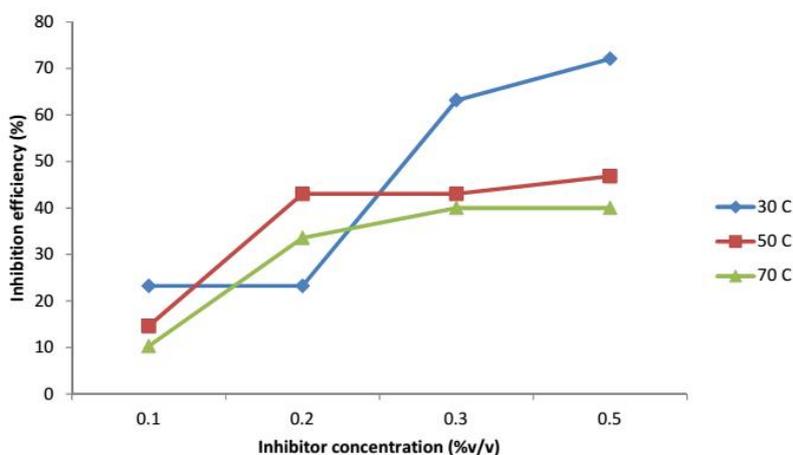
molecules from the alloy surface as a result of dissociation of constituents of inhibiting substance [2, 11].



**Figure 5:** Variation of inhibitor efficiency of guava (*Psidium guajava*) leaf extract on double thermally aged Al-Si-Mg alloy with time of exposure at 50°C.



**Figure 6:** Variation of inhibitor efficiency of guava (*Psidium guajava*) leaf extract on double thermally aged Al-Si-Mg alloy with time of exposure at 70°C.



**Figure 7:** Variation of inhibition efficiency on double thermally aged Al-Si-Mg alloy with concentration of guava (*Psidium guajava*) leaf extract for immersion time of 1hr at different operating temperatures.

### Potentiodynamic polarization

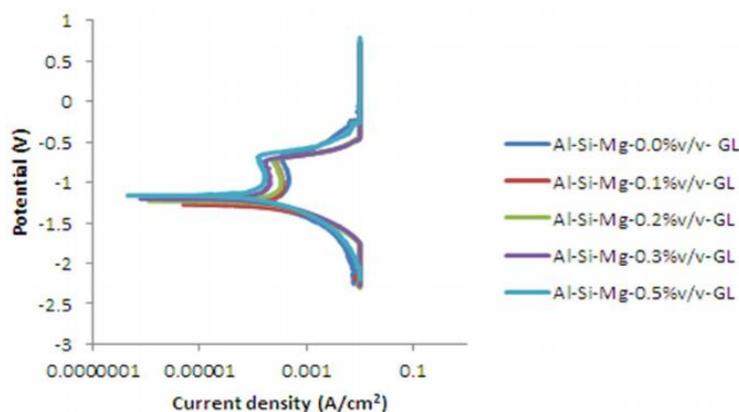
In the potentiodynamic polarization for Al-Si-Mg (SSM-HPDC) alloy in 3.5%wt NaCl/ *Psidium guajava* environment (Table 3), different measurement consisting of potentiodynamic polarization-corrosion rate (PP-CR), potentiodynamic polarization-corrosion density (PP-Jcorr), and linear polarization resistance (LPR) were used as a criteria for evaluation of corrosion resistance of Al-Si-Mg alloy in the environment. Figure 8 indicate the polarization curves for age-hardened Al-Si-Mg (SSM-HPDC) alloy in 3.5%wt NaCl/ *Psidium guajava* at 298K. Generally, the studied environment demonstrated a decreased corrosion rate and corrosion current density with addition of *Psidium guajava* inhibitor. While the corrosion potential ( $E_{\text{corr}}$ ) and polarization resistance ( $R_p$ ) increases with inhibitor concentrations. The corrosion potential of the alloy in blank solution was -1.2771 V but at various concentration of guava leaf extract, the corrosion potential values shifted towards the anodic side (-1.2760 V, -1.2325 V, -1.2070 V, -1.1649 V) indicating that the extract affects the anodic reaction predominantly. This is in agreement with previous reports [1, 12]. The inhibited age-hardened Al-Si-Mg (SSM-HPDC) alloy in 3.5%wt NaCl showed that corrosion rate decreased from 0.936820 mm/yr to 0.478450, 0.447600, 0.176970 and 0.010488 mm/yr at 0.1, 0.2, 0.3, and 0.5%v/v *Psidium guajava*. The linear polarization values showed that addition of guava leaf extract resulted in an increase in polarization resistance value ( $R_p$ ) for the alloy from 195.58 $\Omega$  (uninhibited) to 6736.6 $\Omega$  (inhibited) condition with an IE of 97.1%. The increase in  $R_p$  and corresponding decrease in  $j_{\text{corr}}$  generally suggested an improvement in corrosion resistance of the age-hardened alloy in the presence of inhibitor. This shows that the alloy in both the as-received and age-hardened condition was protected within the immersion time considered [12].

**Table 3.** The potentiodynamic polarization parameters for the corrosion of double thermally aged Al-Si-Mg (SSM-HPDC) in 3.5wt% NaCl in the absence and presence of different concentration of guava leaf extracts.

S/N	Conc. (% v/v)	$E_{\text{corr, obs}}$ (V)	$j_{\text{corr}}$ (A/cm <sup>2</sup> )	Corrosion rate (mm/yr)	Polarization resistance ( $\Omega$ )
12	0.0	-1.2771	8.06E-05	0.936820	195.58
3	0.1	-1.2760	4.12E-05	0.478450	166.87
4	0.2	-1.2325	3.85E-05	0.447600	166.49
5	0.3	-1.2070	1.52E-05	0.176970	383.17
6	0.5	-1.1649	9.03E-07	0.010488	6736.6

### Inhibitor efficiency and adsorption behavior

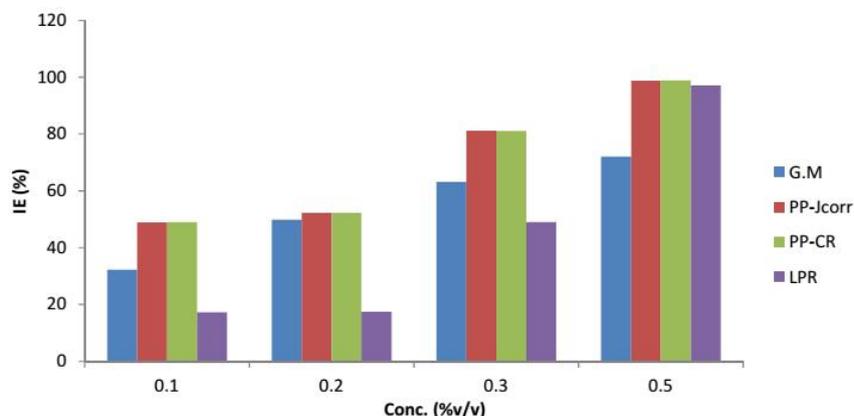
The computed inhibitor efficiency (% IE) of age-hardened Al-Si-Mg (SSM-HPDC) alloy-*Psidium guajava* in 3.5%wt NaCl solution was done using the equation reported [1,12]. The variation in % IE using the gravimetric (G.M), potentiodynamic polarization-corrosion rate (PP-CR), potentiodynamic polarization-corrosion density (PP-Jcorr), and linear polarization resistance (LPR) are presented in Figure 9 for 3.5wt%NaCl/ *Psidium guajava*.



**Figure 8:** Potentiodynamic polarization curves for double thermally aged Al-Si-Mg (SSM-HPDC) alloy in 3.5wt%NaCl in absence and presence of different concentrations of Guava leaf extracts.

**Table 4.** Comparative table of inhibitor efficiency (IE) for age-hardened alloy-3.5wt%NaCl solution / Guava (*Psidium guajava*) leaf extract concentration obtained by gravimetric (G.M), potentiodynamic polarization-corrosion current (PP- $J_{corr}$ ), potentiodynamic polarization-corrosion rate (PP-CR), and linear polarization resistance (LPR).

Conc. (%v/v)	G.M (%)	PP- $J_{corr}$ (%)	PP-CR (%)	LPR (%)
0.0	-	-	-	-
0.1	32.24	48.88	48.92	-17.21
0.2	49.84	52.23	52.22	-17.47
0.3	63.17	81.14	81.11	48.96
0.5	72.10	98.88	98.89	97.10

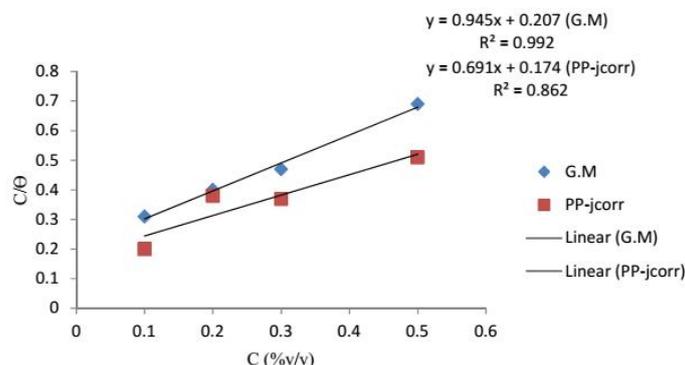


**Figure 9:** Comparative chart of inhibitor efficiency (IE) for double thermally-aged Al-Si-Mg-3.5%wt NaCl solution/ *Psidium guajava* concentration obtained by gravimetric (G.M), potentiodynamic polarization-corrosion rate (PP-CR), potentiodynamic polarization-corrosion current (PP- $J_{corr}$ ) and linear polarization resistance (LPR).

The results (see Table 4) show that addition of *Psidium guajava* increased the IE with an increase in the inhibitor concentrations. This may be attributed to the fact that, as the *Psidium guajava* addition increases, the surface covered by this inhibitor increased hence higher IE was achieved. *Psidium guajava* extracts can be said to exhibit mixed-type corrosion because of the simultaneous change in

the anodic and cathodic region during the electrochemical evaluation of the corrosion condition.

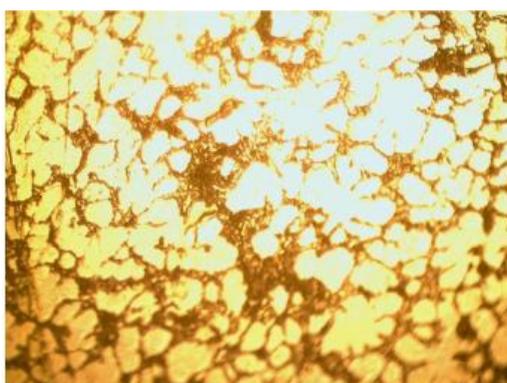
The adsorption mechanism of the extracts in showed that the relationship between  $C/\theta$  and  $C$  is linear for both extracts in simulated seawater environment (Figure 10). Since the linear regression coefficient/or correction factors ( $R^2$ ) close to unity: G.M (0.992), PP (0.862) the adsorption behavior is assumed to have obeyed Langmuir adsorption isotherms. In this case, the adsorption of inhibitor molecules belongs to a monolayer adsorption, assuming that lateral interactions between the inhibitor molecules are ignored.



**Figure 10:** Langmuir isotherm for the adsorption of guava (*Psidium guajava*) leaf extract on double thermally-aged Al-Si-Mg (SSM-HPDC) alloy obtained by gravimetric based-mass loss and potentiodynamic polarization method.

### **Microstructure examination and surface morphology**

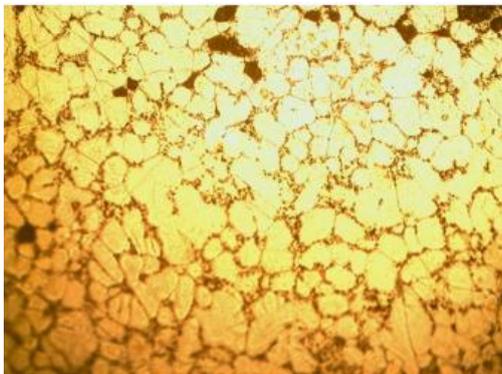
From microstructure analysis, it can be seen that the microstructure of the age-hardened alloy (Plate II) has finer grains and enhanced grain boundaries than the as-received alloy (Plate I). This is based on the fact that there was an increase in volume of the precipitated intermetallic compounds during ageing process and also refinement of the precipitated constituent particles [6].



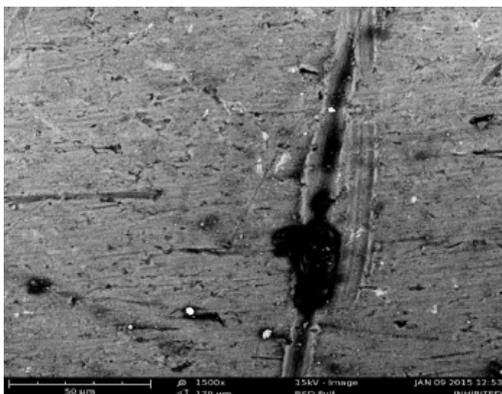
**Plate I:** OPM microstructure of Al-Si-Mg (SSM-HPDC) alloy ( $\times 100$ ). The structure consists of eutectic silicon and intermetallic particles in Al matrix.

The SEM micrographs of uninhibited Al-Si-Mg sample in 3.5%NaCl solution (Plate III) indicate severe surface degradation and pit formation compared to the inhibited sample (Plate IV). This indicates that the guava leaf extract was able to exhibit some degrees of inhibition which retarded the corrosion rate of the Al-Si-

Mg alloy. This occurrence has been attributed to formation of thin oxides on the alloy surface which interfere with the reaction sites, thus impede the formation of pits and their growths.



**Plate II:** OPM microstructure of double thermally-aged Al-Si-Mg (SSM-HPDC) alloy ( $\times 100$ ) showing refinement of precipitated constituent particles with enhanced grain boundaries.



**Plate III:** SEM micrograph of age-hardened Al-Si-Mg (SSM-HPDC) alloy in 3.5wt% NaCl-0.5%v/v guava (*Psidium guajava*) leaf extract after 3 hrs.



**Plate IV:** SEM micrograph of uninhibited age-hardened Al-Si-Mg (SSM-HPDC) alloy in 3.5wt% NaCl after 3 hrs

### Conclusions

The hardness of the Al-Si-Mg (SSM-HPDC) alloy was improved by 28.62% employing the double thermal ageing treatment. *Psidium guajava* have been

demonstrated to be an effective corrosion inhibitor for Al-Si-Mg (SSM-HPDC) alloy in 3.5wt% NaCl. The corrosion resistance and inhibitor efficiency of age-hardened Al-Si-Mg (SSM-HPDC) alloy increased with addition of *Psidium guajava* as inhibitor. IE as high as 72.1% and 98.89% in 3.5wt% NaCl- *Psidium guajava* condition was obtained for the gravimetric and potentiodynamic polarization techniques respectively. The adsorption behavior of *Psidium guajava* leaf extract in 3.5wt% NaCl for age-hardened Al-Si-Mg (SSM-HPDC) alloy can be said to obey the Langmuir adsorption isotherms. Potentiodynamic polarization measurements show that *Psidium guajava* is a mixed type inhibitor predominant to the anodic branches. The SEM examination showed that there was improvement in surface morphology of the as-corroded inhibited alloy compared to uninhibited sample.

### References

1. Abdulwahab M, Popoola API, Fayomi OSI. Int J Electrochem Sci. 2012;7:11706-11717.
2. Asuke F, Yaro S A, Oloche O B, et al . J Met Mat Engineering. 2009;4:48-54.
3. Ayeni FA, Madugu IA, Sukop P, et al . J Min Mat Charact Eng. 2012;11:667-670.
4. Birol Y. J Alloys Compounds. 2009;484:164-167.
5. Hongwei S, Han EH, Lamaka SV, et al. Prog Inorg Coating. 2014;77:765-773.
6. Hossain A, Kurny AS. Int J Chemical, Nuclear, Met Mat Eng. 2014;8:288-293.
7. Howida AF, Tarek MA, Mohamed SE. Int J Electrochem Sci. 2014;9:1565-1582.
8. Loto CA, Loto RT, Popoola API. Int J Phys Sci. 2011;6:2249-2257.
9. Mohammed RA, Abdulwahab M, Madugu IA, et al. J Mater Environ Sci. 2013;4:93-98.
10. Nnanna LA, Obasi VO, Nwadiuku OC, et al. Arch Appl Sci Res. 2012;4:207-217.
11. Olusegun SJ, Adeiza BA, Ikeke KI, et al. J Emerging Trends Eng Appl Sci (JETEAS). 2013;4:138-143.
12. Popoola API, Fayomi OSI, Abdulwahab M. Int J Electrochem Sci. 2012;7:5817-5827.
13. Stephen GY, Sylvester OA, Tersoo GT, et al. Int J Adv Research Chem Sci (IJARCS). 2014;1:38-46.
14. wan Nik WB, Sulaiman O, Ayob AF, et al. Int J Eng Res Appl. 2012;1:723-728.